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# Nano-Engineering of New Photoreactive Materials Having Fluorine Units: Synthesis and Photoalignment of Liquid Crystals

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## ABSTRACT

In order to investigate the photoalignment mechanism of nematic liquid crystals on an oxadiazoles, new 1,3,4-oxadiazole having fluorine atoms as terminal, as well as sideways was synthesized. Linearly polarized UV light shined in these nano-layered films causes the preferential direction of the liquid crystal molecules either perpendicular (in the case of terminal fluoro substitution) or parallel (in the case of sideways fluorine substitution) to the polarized direction. A powerful indication about the relationship between molecular structure and alignment properties where just by changing fluorine atom can change the entire alignment direction is proposed. Such information is indispensable for the design and synthesis of photoalignment materials for liquid crystal displays of high quality.

**Keywords:** Liquid Crystals; Fluorine; Parallel; Perpendicular; Irradiation

## 1. Introduction

Due to high molecular hyperpolarizabilities, elongated nature, chemically stable, sometimes with good electron-transport properties, many of the oxadiazole derivatives displays a number of interesting photo and nonlinear optical effects. Certain oxadiazoles can reveal liquid crystalline features and in particular, can show the elusive biaxial nematic phase. Theoretically proposed in 1970 by Freiser, the biaxiality can be observed in the nematic phase of boomerang-shaped oxadiazoles [1]. Oxadiazole rings can provide not only the lateral dipole from the oxygen and nitrogen atoms, but also the bent shape of the rigid cores. It is due to the polar substituents which are highly variable, and its attaching sites are also adjustable. Therefore, these benefits allow us to access a wide variety of mesophases simply by changing the polar moieties or their attaching positions [2]. Compounds with the 1,3,4-oxadiazole motif are generally chemically stable and have diverse luminescent properties, and are often emissive mesogens with good electron-transport properties [3] which seems to play an important role for the photoalignment properties of these materials. In general, the oxadiazoles are interesting for applications to elec-

troluminescent devices where they are the emissive materials [4].

On the other hand, for good electro-optical performances of liquid crystal displays (LCD), defect free and uniform liquid crystal (LC) alignments are required. Mechanical rubbing of thin polyimide alignment layers with cotton or nylon cloth is the most commonly used alignment method today to achieve uniform LC alignment because this method is highly reproducible and feasible in mass production of LCD [5,6]. However, conventional LCD employs the rubbing technique in order to obtain a uniform LC alignment exhibit asymmetric and narrow viewing angle characteristics and thus their applications to large-size and full color LCD are limited [7]. On the other hand, liquid crystal alignment by photoresponsive materials (nano scale thin photoresponsive solid films) has been studied as an alternative LC alignment method which is a promising technique for the fabrication of LCDs due to its noncontact nature and the possibility of producing a multidomain structure (photo-patterning) [8].

The photoalignment (PhA) effect of the thin films made from photoresponsive materials is due to photo-induced anisotropy of the surface physical properties of these films acting as an LC alignment layer. This surface

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anisotropy is generated due to the selective photoreactions of photoresponsive materials by linearly polarized UV light, for instance. The alignment films made from this kind of materials align the LC molecules along a preferential direction determined by the orientation of the light polarization, which in some cases is along the polarization direction or perpendicular to it. PhA effects of a variety of materials with different photosensitive groups mainly, azobenzenes [9], cinnamates [10], coumarin [11], maleimides [12] etc. were studied extensively during the past two decades [8]. Due to lack of suitable materials and good understanding of the PhA processes, the PhA is still hampered and not able to perform desired results which are demanding for industrial applications in the field of LCDs. Good understanding of the relations between molecular structure and PhA ability of the materials is necessary in order to design and synthesize proper materials for PhA of LCs. Furthermore, depending on the structure of the photoresponsive molecules, the illumination conditions with linear polarized UV light and many other processing parameters, the PhA films may promote orientation of the LC molecules either parallel or perpendicular to the polarization of the incident light with or without pretilt [13]. Owing to its high sensitivity, these kinds of orientation effects are also of great interest for application in nonlinear optics and optical storage devices [6]. It is worth to mention that thin film (typically between 20 - 100 nm) made from photoresponsive materials are useful for many applications [14].

To gain knowledge about the relations between molecular structure and PhA ability, we designed and synthesized in this work heterocyclic compounds having 1,3,4-oxadiazoles bearing 2-fluoro-4-methoxy phenyl moiety and studied their photo alignment properties. The experiments with alignment films prepared from the oxadiazole 4b and oxadiazole 4d reveal different PhA properties closely related to the position of fluorine in their structure.

## 2. Experimental

### 2.1. Materials and Instrumentation

All chemicals were purchased from Aldrich and used without further purification. Melting points were determined by open capillary method and were uncorrected. Elemental analysis was performed on Thermo Fin nigan FLASHEA 1112CHN analyzer. IR spectra (KBr pellets) were recorded on a Shimadzu FT-IR 157 spectrophotometer. <sup>1</sup>H-NMR spectra were recorded on a BRUKER AMX 400MHz spectrometer using TMS as internal standard. Mass spectra were recorded on LCMS-Agilent 1200 series with MSD using 0.1% aqueous TFA in acetonitrile system on C18-BDS column for 10 min duration. Ionization mode is EI for all the compounds.

The purity of the compounds was checked by TLC silica coated plates obtained from Merck. **Figure 1** shows the scheme for the synthesis of oxadiazole molecules whereas **Figure 2** shows the final products obtained from **Figure 1** with different substitutions.

### 2.2. Preparation of Ethyl 2-fluoro-4-methoxybenzoate (2)

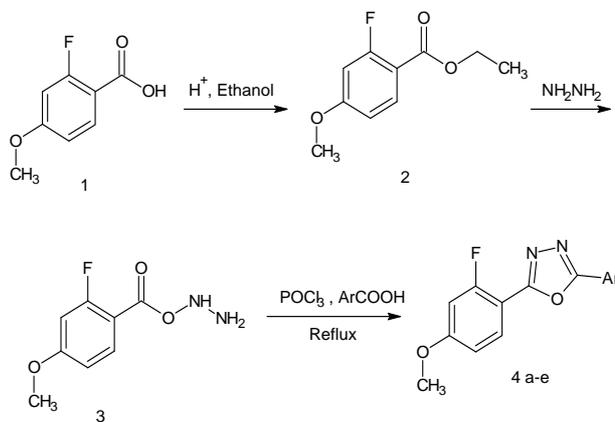
To a mixture of 2-fluoro-4-methoxybenzoic acid (10 g, 0.0587 mol) in ethanol (100 ml) was added conc. sulphuric acid (1 ml) and refluxed for 5 h. The reaction mixture was concentrated, the solid separated was filtered, washed with water and recrystallised with ethanol to give 2 as white crystals. (10 g, 85%) mp. 158°C - 160°C.

### 2.3. Preparation of Ethyl 2-Fluoro-4-methoxybenzohydrazide (3)

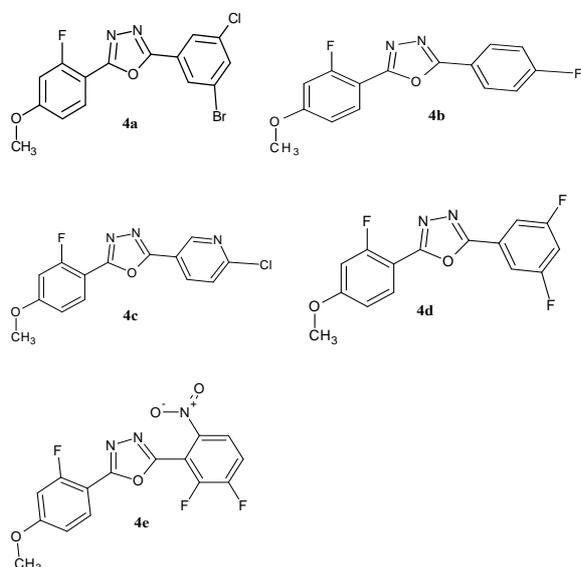
A mixture of ethyl 2-fluoro-4-methoxybenzoate (10 g, 0.0602 mol) and hydrazine hydrate (5.0 ml, 0.1009 mol) in ethanol (100 ml) was heated under reflux for 8 h. The reaction mixture was concentrated and left to cool. The solid product obtained was filtered, washed with water and recrystallised with ethanol to give 3 as colorless crystals. (7 g, 74%) mp. 158°C - 160°C.

### 2.4. General Procedure for Preparation of 2-(2-Fluoro-4-methoxyphenyl)-5-substituted 1,3,4-oxadiazole (4)

An equimolar mixture of acid hydrazide 3 with different aromatic carboxylic acid (a - e) was refluxed with phosphorous oxychloride (10 vol) for 2 - 3 h. The reaction mixture was concentrated through rotovapour, the residue was quenched with ice water and the solid separated was filtered off, washed with water and further purified by recrystallization with ethanol to afford 5-substituted 1,3,4-oxadiazole bearing 2-fluoro-4-methoxy phenyl moiety as white solid [15].



**Figure 1.** Scheme for Synthesis of new oxadiazoles.



**Figure 2.** Chemical structures used in the present investigations.

### 2.5. Preparation of 2-(2-Bromo-5-chlorophenyl)-5-(2-fluoro-4-methoxyphenyl)-1,3,4-oxadiazole (4a)

To a mixture of 2-fluoro-4-methoxybenzohydrazide 3 (1 g, 0.0054 mol) and 2-bromo-5-chloro benzoic acid (1.27 g, 0.0054 mol) was added phosphorous oxychloride (10 ml). The reaction mixture was refluxed at 100°C for 2 h. The reaction mixture was cooled to room temperature, the excess of POCl<sub>3</sub> was concentrated through high vacuum, the residue was quenched with ice and the solid separated was filtered and dried through a pump to afford the title compound 4a as a white solid. (1.9 g, 90%); mp 215°C - 218°C; IR (KBr) cm<sup>-1</sup> 3020 (Ar C-H), C=N (1625), C≡C (1540), C-O (1040, stretch of oxadiazole ring), C-F (1065); mass m/z (M<sup>+</sup>) 382: <sup>1</sup>H NMR (400 MHz-DMSO-d<sub>6</sub>-ppm) δ 8.03 - 8.08 (m, 1H, Ar-H), 8.00 - 8.03 (d, 1H, Ar-H, J = 7.8 Hz), 7.82 - 7.84 (d, 1H, Ar-H, J = 7.0 Hz), 7.38 - 7.43 (m, 1H, Ar-H), 7.12 - 7.18 (dd, 1H, Ar-H, J = 2.4 Hz), 7.00 - 7.04 (dd, 1H, Ar-H, J = 2.31 Hz), 3.88 (s, 3H, -OCH<sub>3</sub>). Anal. Found (calc.) for C<sub>15</sub>H<sub>9</sub>BrClF<sub>2</sub>N<sub>2</sub>O<sub>2</sub> (%): C, 46.97 (47.0); H, 2.36 (2.5); N, 7.30 (7.2).

### 2.6. Preparation of 2-(4-Fluorophenyl)-5-(2-fluoro-4-methoxyphenyl)-1,3,4-oxadiazole (4b)

To a mixture of 2-fluoro-4-methoxybenzohydrazide 3 (1 g, 0.0054 mol) and 4-fluoro benzoic acid (0.76 g, 0.0054 mol) was added phosphorous oxychloride (10 ml). The reaction mixture was refluxed at 100°C for 2 h. The reaction mixture was cooled to room temperature, the excess of POCl<sub>3</sub> was concentrated through high vacuum, the residue was quenched with ice and the solid separated

was filtered and dried through a pump to afford the title compound 4b as a white solid. (1.4 g, 90%); mp 210°C - 212°C; IR (KBr) cm<sup>-1</sup> 3030 (Ar CH), C=N (1635), C≡C (1530), C-O (1030, stretch of oxadiazole ring), C-F (1060); mass m/z (M<sup>+</sup>) 289: <sup>1</sup>H NMR (300 MHz-DMSO d<sub>6</sub>-ppm). δ 8.07 - 8.16 (m, 3H, Ar-H), 7.44 - 7.50 (m, 2H, Ar-H), 7.11 - 7.16 (dd, 1H, Ar-H, J = 2.4 Hz), 7.0 - 7.04 (dd, 1H, Ar-H, J = 2.4 Hz), 3.87 (s, 3H, -OCH<sub>3</sub>). Anal. Found (calc.) for C<sub>15</sub>H<sub>10</sub>F<sub>2</sub>N<sub>2</sub>O<sub>2</sub> (%): C, 61.65 (61.3); H, 3.66 (3.48); N, 9.72 (9.8).

### 2.7. Preparation of 2-Chloro-5-[5-(2-fluoro-4-methoxyphenyl)-1,3,4-oxadiazol-2-yl] Pyridine (4c)

To a mixture of 2-fluoro-4-methoxybenzohydrazide 3 (1 g, 0.0054 mol) and 6-chloronicotinic acid (0.84 g, 0.0054 mol) was added phosphorous oxychloride (10 ml). The reaction mixture was refluxed at 100°C for 2 h. The reaction mixture was cooled to room temperature, the excess of POCl<sub>3</sub> was concentrated through high vacuum, the residue was quenched with ice and the solid separated was filtered and dried through a pump to afford the title compound 4c as a white solid. (1.5 g, 90%); mp 224°C - 235°C; IR (KBr) cm<sup>-1</sup> 3030 (Ar C-H), C=N (1625), C≡C (1560), C-O (1080, stretch of oxadiazole ring), C-F (1040), C-Cl (833); mass m/z (M<sup>+</sup>) 306: <sup>1</sup>H NMR (300 MHz-DMSO-d<sub>6</sub>-ppm) δ 9.08 - 9.09 (d, 1H, Ar-H, J = 2.4 Hz), 8.47 - 8.50 (dd, 1H, Ar-H, J = 1.98 Hz), 8.07 - 8.13 (t, 1H, Ar-H, J = 8.61 Hz), 7.78 - 7.80 (d, 1H, Ar-H, J = 8.4 Hz), 7.12 - 7.17 (dd, 1H, Ar-H, J = 11.2 Hz), 7.01 - 7.05 (dd, 1H, Ar-H, J = 8.73 Hz), 3.88 (s, 3H, -OCH<sub>3</sub>). Anal. Found (calc.) for C<sub>14</sub>H<sub>9</sub>ClF<sub>2</sub>N<sub>3</sub>O<sub>2</sub> (%): C, 55.61 (55.9); H, 3.17 (3.01); N, 13.35 (13.0).

### 2.8. Preparation of 2-(3,5-Difluorophenyl)-5-(2-fluoro-4-methoxyphenyl)-1,3,4-oxadiazole (4d)

To a mixture of 2-fluoro-4-methoxy benzohydrazide 3 (1 g, 0.0054 mol) and 3,5 difluorobenzoic acid (0.85 g, 0.0054 mol) was added phosphorous oxychloride (10 ml). The reaction mixture was refluxed at 100°C for 2 h. The reaction mixture was cooled to room temperature, the excess of POCl<sub>3</sub> was concentrated through high vacuum, the residue was quenched with ice and the solid separated was filtered and dried through a pump to afford the title compound 4d as a white solid. (1.4 g, 87%); mp 265°C - 269°C; IR (KBr) cm<sup>-1</sup> 3070 (Ar CH), C=N (1675), C≡C (1576), C-O (1054, stretch of oxadiazole ring), C-F (1083); mass m/z (M<sup>+</sup>) 307: <sup>1</sup>H NMR (300 MHz-DMSO-d<sub>6</sub>-ppm) δ 8.10 - 8.16 (t, 1H, Ar-H, J = 8.64 Hz), 7.78 - 7.81 (m, 2H, Ar-H), 7.57 - 7.64 (m, 1H, Ar-H) 7.12 - 7.17 (dd, 1H, Ar-H, J = 2.4 Hz), 7.01 - 7.05 (dd, 1H, Ar-H, J = 2.52 Hz) 3.88 (s, 3H, -OCH<sub>3</sub>). Anal. found (calc.) for C<sub>15</sub>H<sub>9</sub>F<sub>3</sub>N<sub>2</sub>O<sub>2</sub>

(%): C, 58.43 (58.8); H, 3.16 (2.96); N, 9.34 (9.2).

## 2.9. Preparation of 2-(2,3-Difluoro-6-nitrophenyl)-5-(2-fluoro-4-methoxyphenyl)-1,3,4-oxadiazole (4e)

To a mixture of 2-fluoro-4-methoxybenzohydrazide 3 (1 g, 0.0054 mol) and 5,6-difluoro-2-nitrobenzoic acid (1.1 g, 0.0054 mol) was added phosphorous oxychloride (10 ml). The reaction mixture was refluxed at 100°C for 2 h. The reaction mixture was cooled to room temperature, the excess of POCl<sub>3</sub> was concentrated through high vacuum, the residue was quenched with ice and the solid separated was filtered and dried through a pump to afford the title compound 4e as a white solid. (1.7 g, 89%); mp 200°C - 203°C; IR (KBr) cm<sup>-1</sup> 3040 (Ar C-H), C=N (1645), C≡C (1540), C-O (1070), stretch of oxadiazole ring, C-F (1070); mass m/z (M<sup>+</sup>) 352; <sup>1</sup>H NMR (400-MHz-DMSO d<sub>6</sub>-ppm). δ 8.59 - 8.63 (dd, 1H, Ar-H, J = 7.2 Hz), 8.33 - 8.38 (dd, 1H, Ar-H, J = 7.2 Hz), 7.99 - 8.04 (t, 1H, Ar-H, J = 7.6 Hz), 7.14 - 7.18 (d, 1H, Ar-H, J = 15 Hz), 7.039 - 7.07 (dd, 1H, Ar-H, J = 2.4 Hz), 3.88 (s, 3H, -OCH<sub>3</sub>). Anal. Found (calc.) for C<sub>15</sub>H<sub>8</sub>F<sub>3</sub>N<sub>3</sub>O<sub>4</sub> (%): C, 51.44 (51.3); H, 2.56(2.4); N, 11.46 (11.3).

## 3. Preparation of Thin Solid Alignment Films

Oxadiazole (1% wt) was dissolved in tetrahydrofuran (THF) solvent and the solution was spin coated onto the surface of previously cleaned glass substrates for 10 seconds at 800 rpm followed by 20 seconds at 2000 rpm. The oxadiazole films were baked for 15 minutes at 100°C. This produces nano surfaces of 1,3,4-oxadiazole films (with thickness of about 50nm) on glass substrates. Experimental quartz substrates also were prepared in a similar way for light absorption study. The glass and quartz substrates covered with photoresponsive film were exposed to linearly polarized UV light using standard USHIO light exposure equipment equipped with deep medium pressure mercury lamp giving 3 mW/cm<sup>2</sup> linearly polarized UV light passing through polarizer and light filter cutting the light below 290 nm. Conventional cells were prepared with the glass substrate after UV Light illumination with a cell gap of 3 - 4 μm. Nematic liquid crystal MLC 6873 - 100 (Δε > 0) was filled at an isotropic phase in the cell via capillary action. The quality of PhA was determined by polarizing optical microscope with crossed polarizers. The direction of the preferred alignment of the liquid crystal molecules was found using the birefringence compensation method.

## 4. Results and Discussion

The formation of oxadiazole 4 was confirmed by re-

cording its IR, NMR & mass spectra. IR spectrum of oxadiazole showed the Ar-H band at 3097 cm<sup>-1</sup>, C=N band at 1594 cm<sup>-1</sup>, C=C band at 1560 cm<sup>-1</sup>, and C-F band at 1093 cm<sup>-1</sup>. The stretch of oxadiazole ring appeared at 1057 cm<sup>-1</sup> due to C-O. Mass spectrum of 4a showed molecular ion peak m/z (M<sup>+</sup>) 364. <sup>1</sup>H-NMR spectrum showed multiplet in the region of δ, 7.00 to 8.07. A singlet appearing at δ, 2.73 is due to three protons of -OCH<sub>3</sub>. Similarly other molecules showed their characteristic peaks [15].

**Table 1** summarizes the results obtained using different substitutions. It is evident from the table that the alignment films made from the compound 4a and 4c, with the substitution of Cl and Br unable to promote alignment of the nematic liquid crystal.

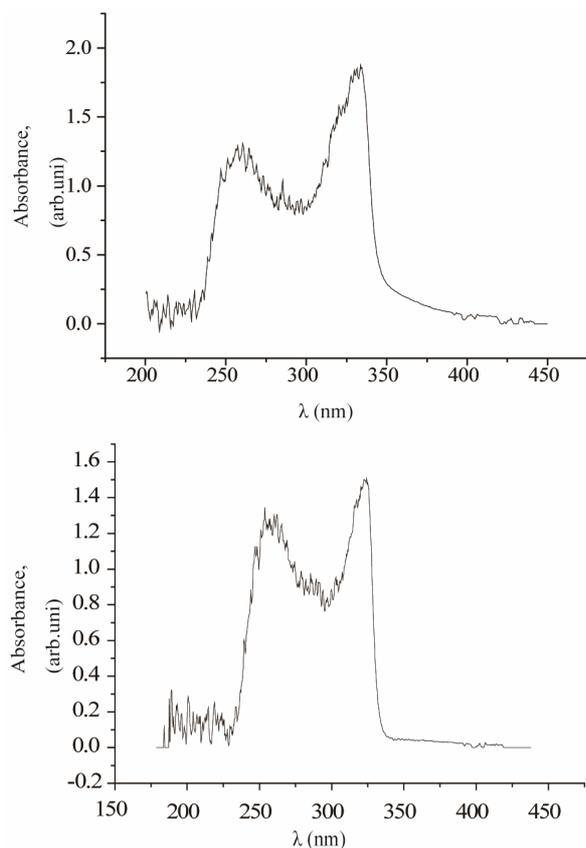
On the other hand, the compound 4b, where fluorine is attached to the end of the molecule and 4d, where fluorine is attached to the both sides of the molecule, was found to produce high quality alignment whereas 4e, with the substitution of two fluorine groups on the same side promoted weak PhA. Therefore, only the experimental results related to compounds 4b and 4d will be discussed here.

UV-vis spectroscopy data of 4b and 4d presented in **Figure 3**, when fluorine is attached to the end of the molecule (in the case of 4b) and fluorine attaches to the both sides of the molecule (in the case of 4d), showed two peaks, one at 260 nm, and the other at ~330 nm. Since peak at 260 nm is non-photoactive, is eliminated by placing 290 nm long pass filter. For the compound comprising of fluorine attached at the end of the molecule (4b), was found to promote liquid crystal planar alignment with preferred direction *perpendicular* to the light polarization direction. Whereas for the compound comprising of fluorine attached to the both sides of the molecules (4d), was found to promote liquid crystal planar alignment with preferred direction *parallel* to the light polarization direction. The quality of the photo alignment was uniform in both cases (see **Figure 4**).

Hence, we may conclude that the position of fluorine plays very vital role in the structure making the liquid crystal molecules to change their direction. This result is

**Table 1. Photoalignment effects and molecular orientation based on the substitution of F, Cl and Br.**

Materials used in this study	Duration of UV exposure (min)	Photoalignment results	Molecular orientation
4a	15	No alignment	No alignment
4b	15	Good	perpendicular
4c	15	No alignment	No alignment
4d	15	Good	parallel
4e	15	Weak	perpendicular



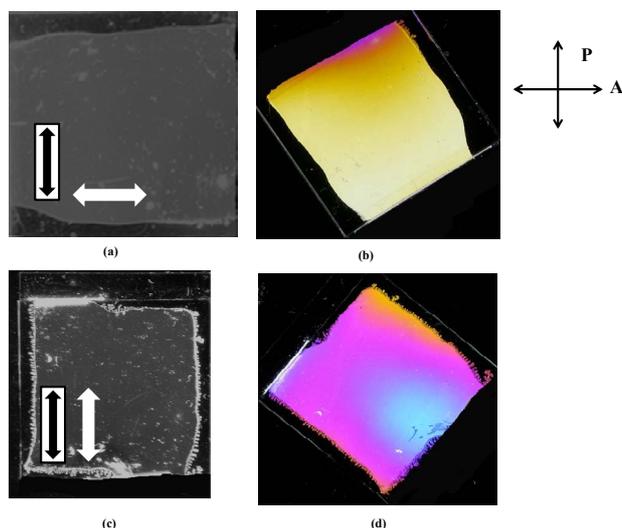
**Figure 3.** UV/vis absorption spectra of the compound 4b (top), 4d (bottom) before illumination. Peak absorbance is around 330 nm. Worthwhile to mention that, after illumination, 4b aligns perpendicular and 4d aligns parallel to the light polarization direction.

very powerful since it helps to design and synthesis of PhA materials of high quality and also controllable polarization directions.

By adding polymer groups in the molecule or by increasing the anisotropy in the structure can increase the stability. As of now present alignment won't stay for long time if we bake the material for a long time. Careful investigations is necessary to improve the stability of the aligning layer.

At wavelength *i.e.*,  $\sim 330$  nm, the absorption wavelength of 4b and 4d compound, the intensity of the medium pressure mercury lamp is low, so it takes much longer time to obtain PhA with this kind of oxadiazoles. However, by using a suitable UV light source, one will be able to get very fast PhA effect using oxadiazoles. On the other hand, one can also increase/decrease the light absorption wavelength of the oxadiazole molecules (say around 313 or 365 nm where UV light source has strong picks) by a proper molecular design which decreases the duration of illumination.

With subsequent ozone treatment of the substrates before spin coating with the solution containing the PhA



**Figure 4.** Photographs of the cells containing alignment layer made from the compound 4b (a and b) and 4d (c and d), respectively, inserted in between two crossed polarizers. The UV light polarization direction is indicated by a black arrow whereas the alignment direction of the liquid crystal in the cells is represented by the white arrow. The dark cell image corresponds to its position between crossed polarizers with the cell optic axis oriented along one of the polarizers whereas the bright image is taken after rotating the cell at  $45^\circ$ . The liquid crystal alignment in the cells is perpendicular, for 4b, and parallel, for 4d, to the light polarization direction exhibiting in both cases excellent quality.

material, improved the quality of alignment promoted by these two compounds drastically.

Let us discuss about the cause of this surprising phenomena. It is reported in the literature that thiophene undergo reorientation upon light illumination [16]. Likewise thiophenes, the polarized UV light drives the molecule of 4b compound to position perpendicular to the light polarization direction since the fluorine is attached laterally to the end of the molecule. This orientation of the molecules of 4b probably corresponds to the minimum molecular excitation by the UV light. For the same reason the molecules of 4d compound, where the position of the fluorine atoms in the structure of 4d is changed  $90^\circ$  with respect to the molecular long axis, orient their molecular axis perpendicular to the light polarization direction. Because of the competing effects of two fluorine atoms, it might be not possible for the entire molecule to change their direction rather they exhibit parallel orientation.

#### 4. Conclusion

We have synthesized new 1,3,4-oxadiazole derivatives, which have been characterized by IR, NMR, mass spectral data and also by C, H, N analyses. Our investigations showed that the molecular structure of the materials possessing PhA ability is of vital importance for the charac-

teristics of the alignment promoted by them. In particular, the position of fluoro substitution in the molecular structure of these materials is important for obtaining PhA of high quality (excellent dark state) with good thermal stability. This study also suggests that the oxadiazoles can successfully be used as PhA materials. Further studies concerning the aligning behaviour are in progress and will be reported soon.

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